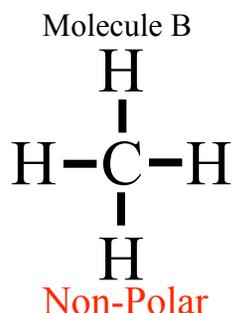
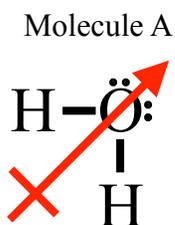
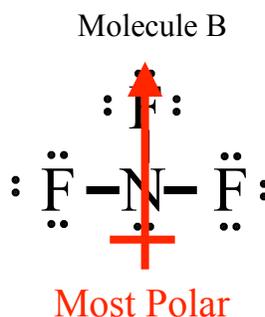
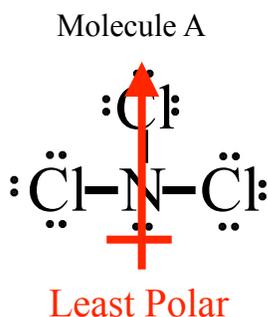


For the following LDMS, draw in the arrow representing the polarity of the molecules. If the molecule is not polar, draw no arrow. Then, answer the questions that follow using A and B to reference the molecules.



- Which molecule will have the higher boiling point? **A**
 Which molecule will have the weaker intermolecular forces? **B**
 Which molecule will have the high volatility? **B**
 Which molecule will more likely be a liquid at room temperature? **A**
 Which one will be a better solvent for polar molecules? **A**
 Which one is least likely to dissolve in water? **B**

For the following LDMS, draw in the arrow representing the polarity of the molecules. Next, determine the polarity of the individual bonds and use these to determine which molecule is most polar. Finally, answer the questions that follow using A and B to reference the molecules.



- Which molecule will have the higher volatility? **A**
 Which molecule will have stronger cohesive forces? **B**
 Which molecule will have the higher boiling point? **B**
 Which molecule will more likely be a gas at room temperature? **A**
 Which one will have stronger intermolecular forces? **B**
 Which molecule will more likely dissolve in water? **B**