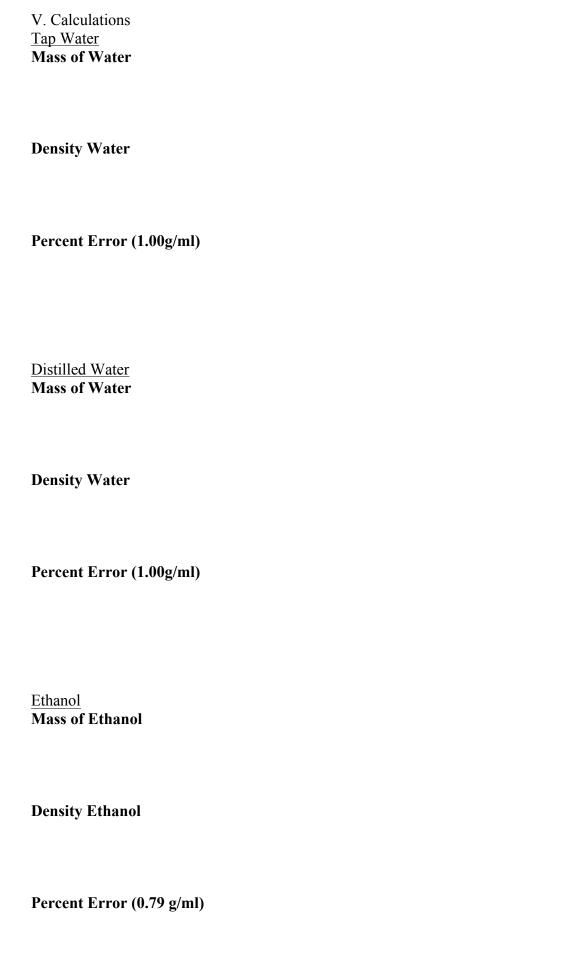
Density Lab
II. Purpose To find the density of solids and liquids and determine the accuracy of their calculated densities.
III. Procedure Tap H ₂ O/Distilled H ₂ O/Ethanol 1. Weigh graduated cylinder 2. Add liquid to graduated cylinder 3. Measure volume of the liquid in the graduated cylinder 4. Weigh graduated cylinder and liquid Zinc/Aluminum 1. Weigh the metal 2. Add water to graduated cylinder 3. Measure the volume of the water 4. Gently add the metal to the water 5. Measure volume of the metal and water
IV. Data (be sure to add units) Tap H ₂ O Mass of graduated cylinder: Mass of graduated cylinder and water: Volume of liquid in graduated cylinder: Distilled H ₂ O Mass of graduated cylinder: Mass of graduated cylinder: Volume of liquid in graduated cylinder: Volume of liquid in graduated cylinder:
Ethanol Mass of graduated cylinder: Mass graduated cylinder and liquid: Volume of liquid in graduated cylinder: Zinc Mass of zinc: Volume of water in graduated cylinder:
Volume of water and zinc in graduated cylinder: Aluminum Mass of aluminum: Volume of water in graduated cylinder: Volume of water and aluminum in graduated cylinder:

Name:__

Per:____



Zinc Volume of Zinc
Density Zinc
Percent Error (From Periodic Table)
Aluminum Volume of Aluminum
Density Aluminum
Percent Error (From Periodic Table)

VI. Conclusions Conceptually and mathematically, what is density?
Why is it that distilled water and tap water have different densities?
Explain the rule for estimating accuracy when making measurements. How did this rule impact the data you recorded with the graduated cylinder and the balance?
Using the concept of density, explain how taking a deep breath will make you float higher in a swimming pool.